

Name \_\_\_\_\_

Date \_\_\_\_\_

## End of Chapter 6 test

*This test and its sample answers have been written by the authors. IB may award marks differently.*

- 1 What is the charge of a sodium ion?  
**A** 1+  
**B** 2+  
**C** 1−  
**D** 2−
- 2 What is the charge of a chloride ion?  
**A** 1+  
**B** 2+  
**C** 1−  
**D** 2−
- 3 What is the charge of an aluminium ion?  
**A** 1+  
**B** 2+  
**C** 3+  
**D** 2−
- 4 What is the charge of a hydroxide ion?  
**A** 1+  
**B** 2+  
**C** 1−  
**D** 2−
- 5 What is the charge of a carbonate ion?  
**A** 1+  
**B** 2+  
**C** 1−  
**D** 2−

- 6 What is the formula of magnesium chloride?
- A  $\text{MgCl}$
  - B  $\text{MgCl}_2$
  - C  $\text{Mg}_2\text{Cl}$
  - D  $\text{Mg}_2\text{Cl}_3$
- 7 What is the formula of sodium oxide?
- A  $\text{NaO}$
  - B  $\text{NaO}_2$
  - C  $\text{Na}_2\text{O}$
  - D  $\text{Na}_2\text{O}_4$
- 8 What is the formula of aluminium oxide?
- A  $\text{AlO}$
  - B  $\text{AlO}_2$
  - C  $\text{Al}_2\text{O}$
  - D  $\text{Al}_2\text{O}_3$
- 9 What is the correct formula for magnesium phosphate?
- A  $\text{Mg}_3(\text{PO}_4)_2$
  - B  $\text{Mg}_2(\text{PO}_4)_3$
  - C  $\text{Mn}_2(\text{PO}_4)_3$
  - D  $\text{MgPO}_4$ .
- 10 The formula for selenium sulfide is  $\text{SeS}_2$ . What would you expect the formula for selenium chloride to be?
- A  $\text{SeCl}$
  - B  $\text{SeCl}_2$
  - C  $\text{SeCl}_3$
  - D  $\text{SeCl}_4$
- 11 Which pair of elements is least likely to form an ionically bonded compound?
- A Li and F
  - B S and O
  - C Mg and O
  - D Zn and Cl

**12** Which substance exhibits only ionic bonding?

- A**  $\text{LiNO}_3$
- B**  $\text{H}_2\text{SO}_4$
- C**  $\text{NH}_4\text{Cl}$
- D**  $\text{CaBr}_2$

**END OF TEST**